

The Hydrogen Bond - fairly strong IMF

(very important Biologically) need 2 things for an H-Bond
 $\sim \frac{1}{20}$ the strength of covalent bond.

DONOR

ACCEPTOR

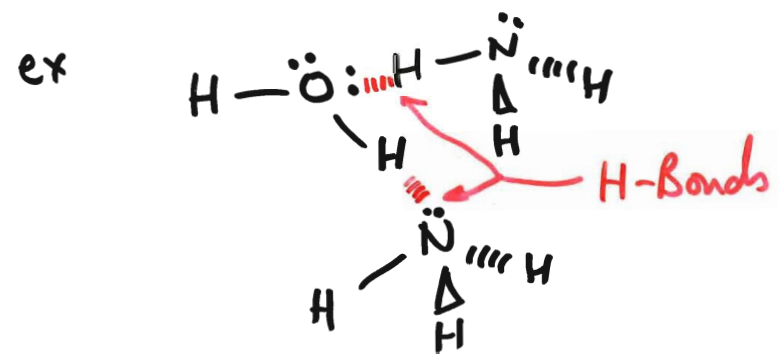
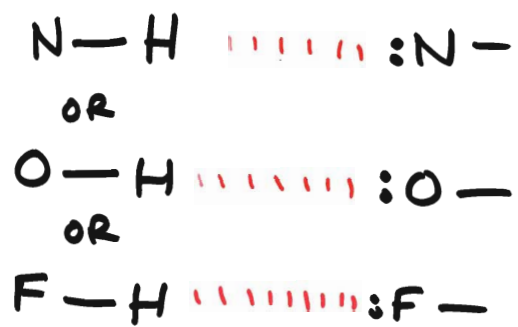
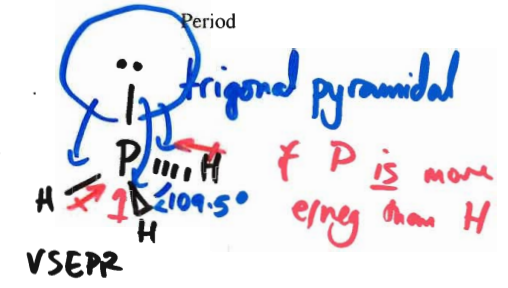
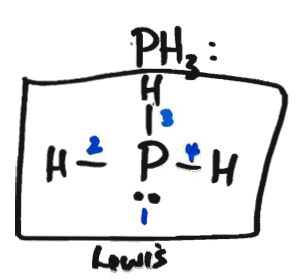
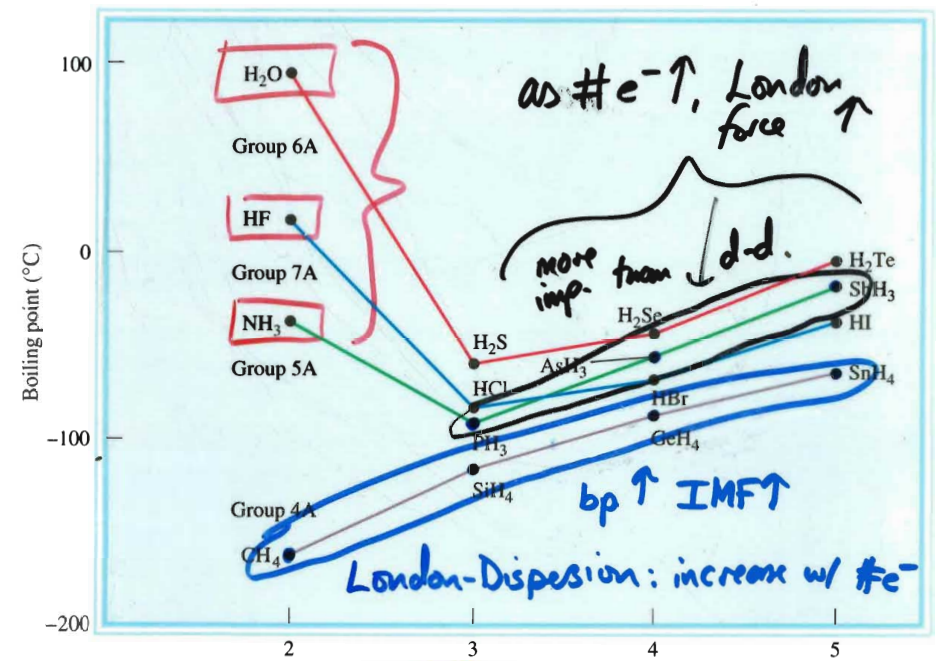
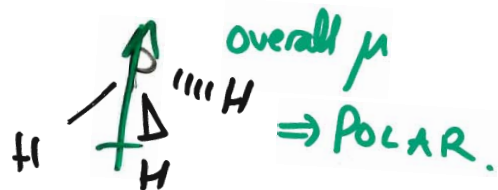


Fig. 12-6

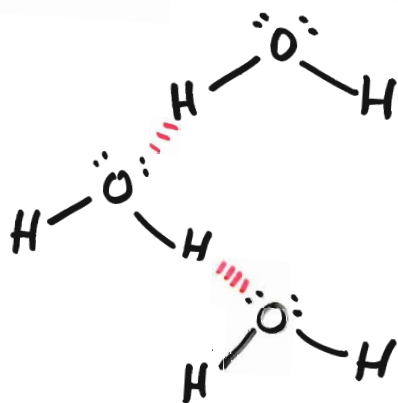
Anomolously high bp
 - from H-Bonding

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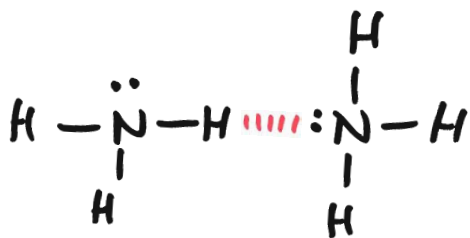




H₂O



NH₃



H-F

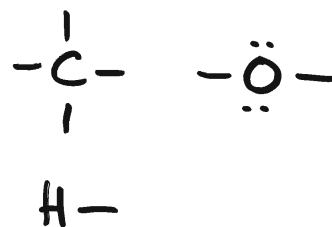


Q. Does H-Bonding exist between molecules of C₂H₆O? If so...

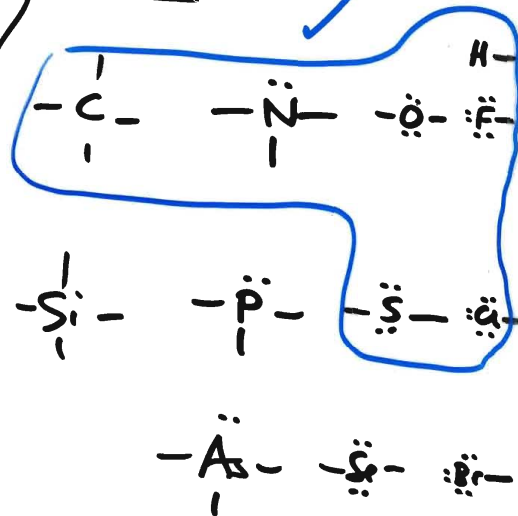
- (1) Show the H-Bonds
- (2) Explain whether bp will be \uparrow or \downarrow .

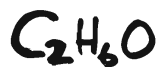
C₂H₆O

Lewis

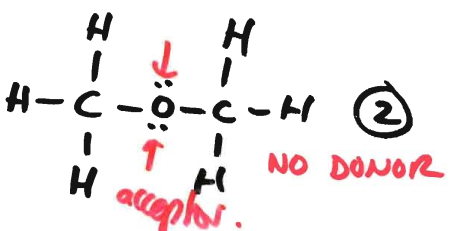
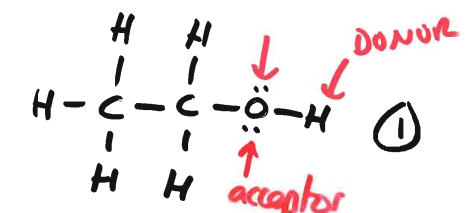
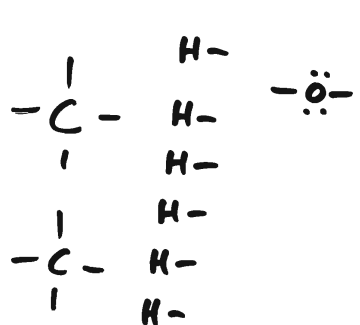


Pattern

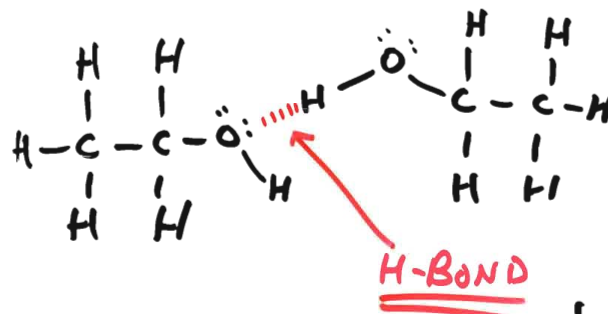
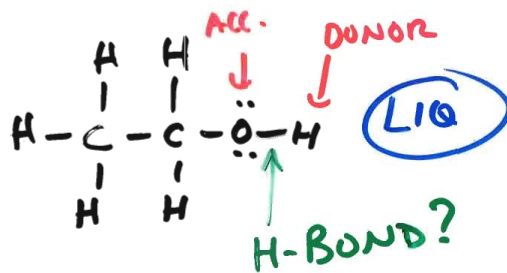
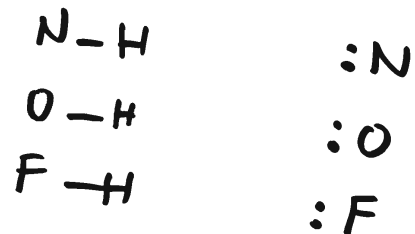




structural isomers.



① H-Bonding
DONOR + ACCEPTOR.



IMF: London-Dispersion, dipole-dipole
H-Bonding } strong IMF

bp ↑

